

Photocatalytic Hydrogen Generation from a Visible-Light-Responsive Metal–Organic Framework System: Stability versus Activity of Molybdenum Sulfide Cocatalysts

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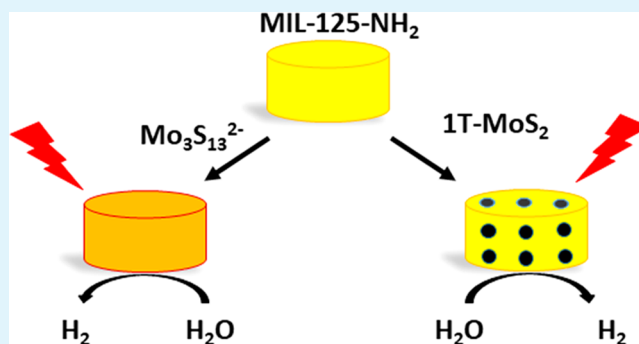
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Supporting Information

ABSTRACT: We report the use of two earth abundant molybdenum sulfide-based cocatalysts, $\text{Mo}_3\text{S}_{13}^{2-}$ clusters and 1T- MoS_2 nanoparticles (NPs), in combination with the visible-light active metal–organic framework (MOF) MIL-125- NH_2 for the photocatalytic generation of hydrogen (H_2) from water splitting. Upon irradiation ($\lambda \geq 420$ nm), the best-performing mixtures of $\text{Mo}_3\text{S}_{13}^{2-}$ /MIL-125- NH_2 and 1T- MoS_2 /MIL-125- NH_2 exhibit high catalytic activity, producing H_2 with evolution rates of 2094 and 1454 $\mu\text{mol h}^{-1} \text{g}_{\text{MOF}}^{-1}$ and apparent quantum yields of 11.0 and 5.8% at 450 nm, respectively, which are among the highest values reported to date for visible-light-driven photocatalysis with MOFs. The high performance of $\text{Mo}_3\text{S}_{13}^{2-}$ can be attributed to the good contact between these clusters and the MOF and the large number of catalytically active sites, while the high activity of 1T- MoS_2 NPs is due to their high electrical conductivity leading to fast electron transfer processes. Recycling experiments revealed that although the $\text{Mo}_3\text{S}_{13}^{2-}$ /MIL-125- NH_2 slowly loses its activity, the 1T- MoS_2 /MIL-125- NH_2 retains its activity for at least 72 h. This work indicates that earth-abundant compounds can be stable and highly catalytically active for photocatalytic water splitting, and should be considered as promising cocatalysts with new MOFs besides the traditional noble metal NPs.

KEYWORDS: metal–organic framework, hydrogen, molybdenum sulfide, photocatalysis, visible light



Metal–organic frameworks (MOFs) are crystalline porous materials prepared by the self-assembly of metal ions or clusters with organic ligands to form 1-, 2-, or 3-dimensional structures. Because of their high porosity and structural tunability, MOFs are promising for applications in many areas, such as gas storage and separation,¹ catalysis,² and sensing,³ among others. In recent years, several semiconducting MOFs have also been employed as photocatalysts for photocatalytic reactions since the use of MOFs can overcome the disadvantages of the traditional photocatalysts such as the negligible absorptivity of visible light of TiO_2 or the low water stability of CdS. Among photocatalytic reactions, the generation of hydrogen (H_2) from water is of great interest mainly because of the potential of H_2 as a clean energy carrier for a sustainable energy future.⁴ To efficiently produce H_2 , MOFs are often coupled with a small amount (<10 wt %) of a cocatalyst, for example, platinum (Pt) nanoparticles (NPs), that attract the photogenerated electrons, and a sacrificial

electron donor which donates its electrons to the photo-generated holes to protect the MOF from degradation.

Within the MOF field, the main research efforts in recent years have been focused on the preparation of new visible-light responsive MOFs since the practical use of these MOFs for photocatalytic H_2 generation would require the absorption of sunlight, 44% energy of which is provided by light in the visible domain. Up to now, several transition metal-based MOFs, such as MIL-125- NH_2 , UiO-66- NH_2 , Al-PMOF, and others, have been known to produce a fairly high amount of H_2 under visible light irradiation when they are combined with Pt NPs.^{5–8} Given the success of Pt NPs as a cocatalyst in H_2 production for a large range of different photoactive materials,

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it has become the “gold” standard in the field. Interestingly, there are also a few reports indicating that suitable earth-abundant cocatalysts can promote H₂ generation,^{9,10} and some of them, such as a Co complex or Ni₂P NPs, can outperform Pt NPs with the same MOF in terms of H₂ evolution rate and apparent quantum yields.^{11–13} The fact that a careful choice of the cocatalyst can significantly enhance the H₂ generation performance adds a new dimension to this system since the number of possible earth-abundant complexes and NPs with oxides, sulfides, phosphides, or nitrides is essentially limitless. In this work, we investigate the efficiency of molybdenum sulfide-based cocatalysts when combined with MOFs.

Natural molybdenum sulfide (2H-MoS₂, H stands for hexagonal symmetry), which is an indirect-band gap (1.29 eV) semiconductor, is an earth-abundant compound that has been well-studied due to its unique electronic, optical and magnetic properties, as well as its potential for applications in electronics, optoelectronics, energy devices and in H₂ generation.¹⁴ The sulfur edge sites and vacancy defects of 2H-MoS₂ were found to be the active sites for the H₂ generation reaction,^{15,16} while the basal planes are catalytically inert. Reducing the size of the 2H-MoS₂ particles to achieve greater density of edge sites has shown an improved photocatalytic H₂ evolution activity.¹⁷ Interestingly, this contraction can even be accomplished at the molecular level; one example is the Mo₃S₁₃²⁻ cluster, (Figure 1a), which

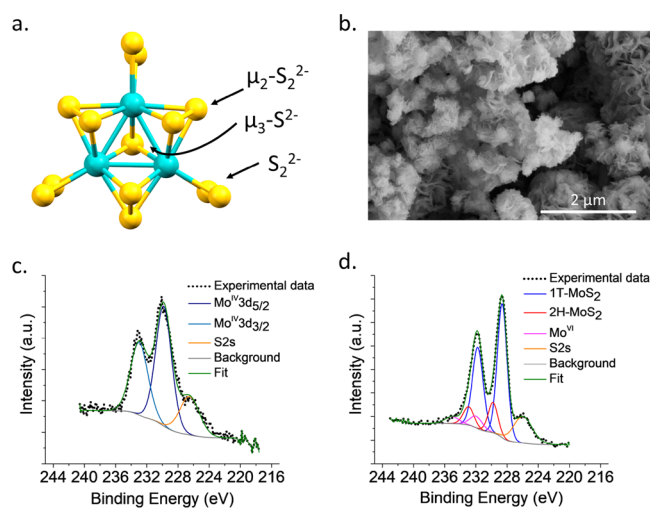


Figure 1. (a) X-ray crystal structure of the Mo₃S₁₃²⁻ cluster (color code: turquoise Mo, yellow S).¹³ (b) SEM image of the 1T-MoS₂. (c and d) XPS spectra of Mo₃S₁₃²⁻ and 1T-MoS₂, respectively.

essentially resembles the smallest fraction of MoS₂. Another strategy to increase the catalytic activity of MoS₂ is to tweak its 2H crystal phase to obtain a polymorph with increased electrical conductivity and hence enhance the electron transfer to the active sites. For example, the 1T-MoS₂ (T stands for trigonal symmetry) can be achieved by a lithium (Li) intercalation-exfoliation procedure resulting in a metastable metallic phase.¹⁸ Despite their promising catalytic activity for H₂ generation,^{19,20} both Mo₃S₁₃²⁻ and 1T-MoS₂ have not been employed as cocatalysts with MOFs in photocatalytic water reduction.

Herein, we report the combination of Mo₃S₁₃²⁻ and 1T-MoS₂ with the well-known visible-light-responsive Ti-based MIL-125-NH₂ for photocatalytic H₂ generation. Both systems

exhibit excellent evolution rates and apparent quantum yields; in addition, the easy synthesis and durable performance of the 1T-MoS₂ NPs suggest that not only can earth-abundant NPs be good cocatalysts but also they can be considered as common cocatalysts, besides the standard Pt NPs, for photocatalytic testing of novel MOF photocatalysts.

The Mo₃S₁₃²⁻ clusters (full chemical formula (NH₄)₂Mo₃S₁₃·2H₂O) and the 1T-MoS₂ NPs were synthesized based on slight modifications of the reported procedures (SI). The cluster contains three Mo^{IV} ions and three different types of S ligands, including three bridging μ_2 -S₂²⁻, three terminal S₂²⁻, and one μ_3 -S₂²⁻. Except for the μ_3 -S₂²⁻, all the other sulfur atoms are appropriately positioned for effective catalysis.²¹ The good match of the powder X-ray diffraction (PXRD) patterns (Figure S1) and the presence of the X-ray photoelectron spectroscopy (XPS) Mo^{IV}3d_{5/2} and Mo^{IV}3d_{3/2} peaks at 229.84 and 232.97 eV (Figure 1c), respectively, confirmed that the Mo₃S₁₃²⁻ clusters were formed without oxidized impurities. Immersing the crystalline powder of Mo₃S₁₃²⁻ in water for 24 h did not affect its crystallinity (Figure S1). Cyclic voltammetry studies were performed in a dimethylformamide (DMF) solution of Mo₃S₁₃²⁻; the cyclic voltammogram (Figure S5) displays the first Mo^{IV} → Mo^{III} reduction peak at \sim -0.44 V, which is less negative than the energy of the conduction band of MIL-125-NH₂ (\sim -0.60 V),²² suggesting that Mo₃S₁₃²⁻ can be used as a cocatalyst with MIL-125-NH₂ for the reduction of water (Figure S6).

The 1T phase of MoS₂ has often been obtained via a Li intercalation-exfoliation process, which involves soaking the bulk 2H-MoS₂ material in hexane solution of *n*-butyl lithium under inert atmosphere for several days followed by exfoliation using ultrasonication.¹⁸ Because of the low yield of this procedure, we prepared our 1T-MoS₂ following a recently reported one-step hydrothermal method with a slight modification (SI) that can easily give a high amount of the product.²³ The PXRD pattern of the 1T-MoS₂ sample (Figure S2) displays comparable Bragg reflections with those of the commercial MoS₂; however, the peaks are broader due to their nanoscale size. Immersing the powder in water for 24 h did not cause any changes in the pattern (Figure S2). Scanning electron microscopy (SEM) images of the 1T-MoS₂ (Figure 1b) show that the sample contains aggregates of flakes, with the size varying between few hundred nm to \sim 2 μ m. As can be seen in the deconvoluted XPS spectrum in Figure 1d, the majority of the sample contains 1T-MoS₂ (\sim 75%, Mo^{IV}3d_{5/2} and Mo^{IV}3d_{3/2} peaks at 228.62 and 231.75 eV, respectively), \sim 17% of the sample is the 2H-MoS₂ (Mo^{IV}3d_{5/2} and Mo^{IV}3d_{3/2} peaks at 229.80 and 232.93 eV, respectively), while the oxidized impurities containing Mo^{VI} comprise \sim 8% (Mo^{VI}3d_{5/2} and Mo^{VI}3d_{3/2} peaks at 231.94 and 235.07 eV, respectively). Raman spectrum of the 1T-MoS₂ sample (Figure S12) shows peaks at \sim 150 cm⁻¹ (J₁), \sim 230 cm⁻¹ (J₂), and \sim 334 cm⁻¹ (J₃), along with peaks at \sim 280 cm⁻¹ (E_{1g}), \sim 376 cm⁻¹ (E_{1g}), \sim 404 cm⁻¹ (A_{1g}), which are consistent with the ones reported for Li-intercalated 1T-MoS₂.¹⁴

Both crystalline powders of Mo₃S₁₃²⁻ and 1T-MoS₂ were subjected to photocatalytic testing with MIL-125-NH₂ by simply mixing them in a 25.0 mL vial containing 17.0 mL solution of acetonitrile (MeCN): triethylamine (TEA): water (H₂O) (79.0:16.1:4.9 v/v/v) mixture. The solution was purged with N₂ before any test. Under visible light irradiation (xenon lamp, long-pass cutoff filter $\lambda \geq$ 420 nm) and by varying the added amount of Mo₃S₁₃²⁻ and 1T-MoS₂ while keeping the

amount of MIL-125-NH₂ constant (17.0 mg), it was found that with 0.82 wt % added amount of the cocatalyst in both cases, the photocatalytic systems produced their highest H₂ evolution rates (Figures S16 and S17), which are 2094 and 1454 μmol h⁻¹ g⁻¹ MOF for Mo₃S₁₃²⁻/MIL-125-NH₂ and 1T-MoS₂/MIL-125-NH₂, respectively. These values are much higher than the H₂ evolution rate of 68 μmol h⁻¹ g⁻¹ MOF when the same added amount of the commercial MoS₂ (Aldrich, <2 μm) was used as the cocatalyst, which exemplifies the enhanced catalytic activity of MoS₂ by either decreasing the size of its particles or increasing their electrical conductivity. These H₂ evolution rates are also much higher than those of other photocatalytic systems with MIL-125-NH₂, for example, the H₂ evolution rates of Pt/MIL-125-NH₂, Co-oxime@MIL-125-NH₂, and Ni₂P/MIL-125-NH₂ in similar reaction conditions are 269, 637, and 894 μmol h⁻¹ g⁻¹ MOF.^{11,13} In fact, to the best of our knowledge, these rates are the highest values reported to date for visible-light-driven photocatalysis with MOFs incorporated with a cocatalyst. It is worth noting that several (semi)conductors such as ZnIn₂S₄ or graphene were employed to form composites with MOFs to significantly enhance their H₂ evolution rates but they are present in high weight percentages (>50 wt %) and are not considered cocatalysts.^{24,25}

Apparent quantum yields were measured for the best-performing mixtures of Mo₃S₁₃²⁻/MIL-125-NH₂ and 1T-MoS₂/MIL-125-NH₂. The measurement were based on the ferrioxalate actinometry method, which first involves the determination of the incident radiation over time (number of photons in a unit of time) of the xenon lamp. The apparent quantum yield (Φ) is then calculated based on eq 1

$$\Phi(\%) = \frac{2 \times \frac{n_{\text{H}_2}}{t_1}}{\frac{n_{\text{photon}}}{t_2}} \times 100\% \quad (1)$$

where n_{H_2} is the number of moles of H₂ produced in the time t_1 , n_{photon} is the number of moles of photons irradiated during the time t_2 . The apparent quantum yields for Mo₃S₁₃²⁻/MIL-125-NH₂ and 1T-MoS₂/MIL-125-NH₂ are 11.0 and 5.8% at 450 nm, respectively, which are excellent compared to the other reported photocatalytic MOF-based systems.^{11,13,26–29}

It is worth noting that the performance of a cocatalyst/MOF system depends on several factors as follows: (i) the driving force for electron transfer from the MOF to the cocatalyst, which depends on the energy gap between them, (ii) the contact/interaction between the MOF and the cocatalyst, and (iii) the inherent catalytic activity of the cocatalyst. To investigate the electron transfer from MIL-125-NH₂ to Mo₃S₁₃²⁻ and 1T-MoS₂ in their suspension mixtures, luminescence emission ($\lambda_{\text{ex}} = 420$ nm) of these solutions were measured together with the emission by MIL-125-NH₂ itself. As shown in Figure 2a, the presence of the cocatalysts (0.82 wt %) quenches the luminescence of the MOF since some of the photogenerated electrons are nonradiatively transferred to the Mo₃S₁₃²⁻ clusters or 1T-MoS₂ NPs. It appears that the Mo₃S₁₃²⁻ clusters do not attract electrons as efficient as the 1T-MoS₂ NPs; however, since the clusters are relatively soluble in the MeCN/TEA/H₂O solution and are adsorbed by the MIL-125-NH₂ (Figure 2b), the interaction between each cluster and the MOF is maximized, together with the inherent high catalytic activity, which apparently compensates for their lower electron affinity. Density func-

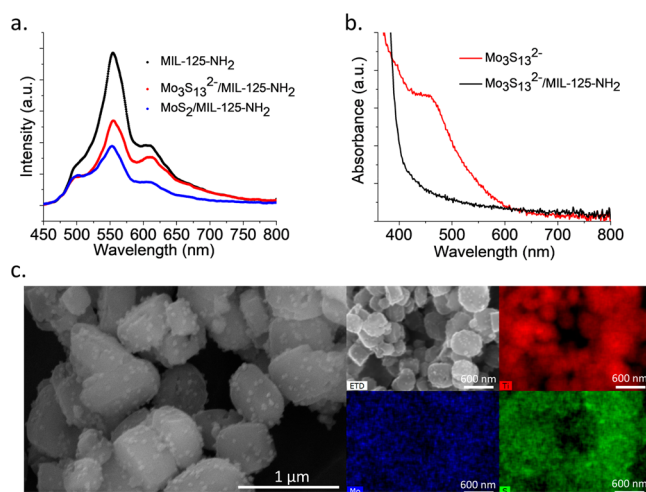


Figure 2. (a) Luminescence emission spectra ($\lambda_{\text{ex}} = 420$ nm) of MIL-125-NH₂ without and with the cocatalysts. (b) UV-vis absorption spectra of the MeCN/TEA/H₂O solution of Mo₃S₁₃²⁻ and of the supernatant after adding MIL-125-NH₂, showing the disappearance of the absorption peak at ~480 nm, indicating that the clusters are adsorbed by the MOF. (c) SEM and EDX mapping images of 1T-MoS₂/MIL-125-NH₂ after 72 h of irradiation, with the 1T-MoS₂ NPs evenly embedded on the surface of the MIL-125-NH₂ crystals.

tional theory (DFT) calculations (SI) were performed and illustrated that due to the relatively large size (6.5 Å) of the Mo₃S₁₃²⁻ cluster, the probability for its diffusion from one pore to another is negligible, suggesting that the adsorption only occurs on the MOF's surface. Nevertheless, the energy-dispersive X-ray (EDX) mapping images (Figure S11) of Mo₃S₁₃²⁻/MIL-125-NH₂ powder after 8 h of irradiation confirmed the even distributions of Mo₃S₁₃²⁻ on the external surface of MIL-125-NH₂ crystals.

SEM and EDX mapping images were also obtained for 1T-MoS₂/MIL-125-NH₂ (Figure 2c). As can be clearly seen in the SEM image, the NPs attach on and cover the surface of the MIL-125-NH₂ crystals allowing for the enhanced interaction between them. It was reported that the basal plane of 1T-MoS₂ nanosheets is also active for the H₂ generation reaction.¹⁴ Therefore, for 1T-MoS₂, their high electron affinity, their relatively good contact with the MOF and their high catalytic activity are combined to give the observed high photocatalytic performance.

For practical applications, an effective cocatalyst should maintain its activity over a long period of time. Recycling experiments were therefore conducted, in which after every 12 h the reactor containing the best-performing cocatalyst/MIL-125-NH₂ was purged with N₂ gas to release the H₂ produced in the head space before a new cycle. While the 1T-MoS₂/MIL-125-NH₂ maintained its H₂ evolution rate (Figure 3b), the Mo₃S₁₃²⁻/MIL-125-NH₂ slowly lost its performance (Figure 3a), probably due to the gradual decomposition of the clusters over time. The low stability of Mo₃S₁₃²⁻ in basic solutions was also previously observed.¹⁹ PXRD patterns of Mo₃S₁₃²⁻/MIL-125-NH₂ and 1T-MoS₂/MIL-125-NH₂ after irradiation were collected; however, the peaks for Mo₃S₁₃²⁻ and 1T-MoS₂ were not observed due to the dissolution of the former and the very low concentration of the latter. Nevertheless, the MIL-125-NH₂, being the main component of the photocatalytic system, was stable as confirmed by the retention of its crystallinity in the PXRD pattern (Figures S3

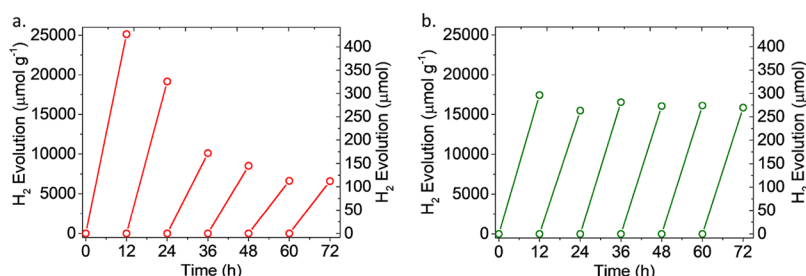


Figure 3. H₂ generation from (a) Mo₃S₁₃²⁻/MIL-125-NH₂ and (b) 1T-MoS₂/MIL-125-NH₂ after 6 cycles (12 h for each cycle).

and S4) and its high porosity even after 72 h of irradiation (Figure S14). This experiment highlights that while the molecular cocatalysts (Mo₃S₁₃²⁻, Co-oxime,¹¹ and others^{12,30}) can have a very high activity, their stability over time is questionable and more robust molecular catalysts are needed for this reaction. On the other hand, although the NPs might have slightly lower catalytic performance, their durability is superior to the molecular ones.

In conclusion, we reported the synthesis and characterization of two molybdenum sulfide-based cocatalysts and studied their performance when combined with MIL-125-NH₂ for visible-light driven photocatalytic H₂ generation. The 1T-MoS₂ not only possess high catalytic activity (the H₂ evolution rate of 1T-MoS₂/MIL-125-NH₂ is only lower than the one of Mo₃S₁₃²⁻/MIL-125-NH₂ and significantly higher than any other reported values for visible-light responsive MOFs incorporated with a cocatalyst), but also maintains its performance for a long period of time. Since almost all recent new MOFs have been mostly combined with Pt NPs for photocatalytic testing, we believe that the easily synthesized 1T-MoS₂ NPs can also be employed as one of the common cocatalysts for these prospective reactions.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acsami.8b10010.

Experimental methods, synthetic procedures, and characterizations (PDF)

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Notes

The authors declare the following competing financial interest(s): We filed a patent application with some of the results presented in this manuscript.

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