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2D Materials



PAPER

Large-area MoS₂ grown using H₂S as the sulphur source

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Abstract

We report on the growth of molybdenum disulphide (MoS₂) using H₂S as a gas-phase sulfur precursor that allows controlling the domain growth direction of domains in both vertical (perpendicular to the substrate plane) and horizontal (within the substrate plane), depending on the H₂S:H₂ ratio in the reaction gas mixture and temperature at which they are introduced during growth. Optical and atomic force microscopy measurements on horizontal MoS₂ demonstrate the formation of monolayer triangular-shape domains that merge into a continuous film. Scanning transmission electron microscopy of monolayer MoS₂ shows a regular atomic structure with a hexagonal symmetry. Raman and photoluminescence spectra confirm the monolayer thickness of the material. Field-effect transistors fabricated on MoS₂ domains that are transferred onto Si/SiO₂ substrates show a mobility similar to previously reported exfoliated and chemical vapor deposition-grown materials.

Main text

Molybdenum disulfide (MoS₂) has attracted extensive attention for a variety of next generation electrical and optoelectronic device applications because of its unique two-dimensional (2D) nature. In the bulk, this material has a crystalline structure consisting of covalently bonded layers weakly coupled to each other by weak van der Waals forces. Bulk MoS₂ is a semiconductor with an indirect bandgap of 1.2 eV [1], whereas monolayer MoS₂ shows a direct gap of 1.8 eV [2–4] due to the 2D confinement [5]. Using scotch tape [6] or liquid-phase exfoliation [7, 8] monolayer MoS₂ can be easily obtained. Versatile devices including field effect transistors with high on/off current ratio [9], memory cells [10], ultrasensitive photodetectors [11, 12] and nanopores [13] were demonstrated on 2D crystals of MoS₂ obtained by mechanical exfoliation from bulk samples.

Nevertheless, exfoliation is not scalable for large-scale device fabrication resulting from the absence of controllable thickness, size and uniformity of the 2D film. Because of this, various chemical synthesis

approaches focused on growth of MoS₂ have been developed [14–18]. Among them, chemical vapor deposition (CVD) is the most promising method to synthesize monolayer MoS₂ from triangular islands up to hundreds of micrometers in size to a large-scale film [17, 18]. Moreover, the proper selection of growth substrate and its pretreatment enables the growth of high-quality centimeter-scale continuous monolayer MoS₂ film that is formed from merging single-crystalline domains with the majority of them having the same lattice orientation [19]. In addition to recent reports on monolayer growth, polycrystalline thin films of transition metal dichalcogenides have been prepared by different methods in the past, such as sputtering [20, 21], sulfurization/selenization of transition metal [22–24] or WO₃ films [25, 26], sulfurization of thin tungsten films using H₂S [27] and sputtering of tungsten in an Ar/H₂S atmosphere [28]. Depending on the growth conditions, two types of crystal orientations were obtained, with the basal plane of the crystallites being predominantly oriented either in parallel or perpendicular directions relative to the substrate surface. Recently, a rapid sulfurization/

selenization process to convert Mo and W thin films into polycrystalline molybdenum dichalcogenide films was also demonstrated, where sulfur/selenium powders as the precursors and Ar as a carrier gas were used [29–31]. Due to chemically reactive edge sites, such structures are useful as an active material for catalytic reactions [32, 33].

In addition to the most commonly used CVD methods that involve a solid sulfur source, a highly reactive sulfuric precursor in the form of H_2S has been used for wafer-scale growth of MoS_2 films on different substrates [34–36]. These previously reported growth methods are however unable to produce monolayer MoS_2 due to the relatively large thickness of molybdenum films pre-deposited on insulating substrates and used as the Mo source [34, 35]. Gold-assisted CVD [36], in which an Mo–Au surface alloy is formed during the reaction of $\text{Mo}(\text{CO})_6$ with Au thin film, is also unable to produce monolayer MoS_2 . Here, we report on the growth of MoS_2 on sapphire substrates using a chemical reaction between MoO_3 and $\text{H}_2\text{S}/\text{H}_2$ gas mixture as a gas-phase sulfur precursor. Controlling the ratio of $\text{H}_2\text{S}:\text{H}_2$ in the reaction gas mixture and the temperature during growth at which they are injected allow us to regulate the growth direction of domains and switch between multilayer growth in the vertical direction or monolayer growth in the horizontal direction. We find that the injection temperature of the $\text{H}_2\text{S}/\text{H}_2$ reaction gas mixture is the critical factor for controlling the growth direction, and observe a transition from vertical to horizontal growth modes with decreasing injection temperature. The optical and electrical properties of the horizontal MoS_2 are comparable to those obtained using other growth methods, allowing its use in future electronic and optoelectronic devices.

Growth procedure

Multilayer vertical and monolayer MoS_2 has been grown by CVD on *c*-plane sapphire. Prior to growth, sapphire substrates were consecutively cleaned by acetone/isopropanol/DI-water and then annealed for 1 h at 1000 °C. For growth, they were placed face-down above a crucible containing ~5 mg MoO_3 ($\geq 99.998\%$ Alfa Aesar) and loaded into a 2 inch split-tube furnace. The temperature is ramped up to 700 °C with a rate of 50 °C min^{-1} and maintained at this target temperature for 10 min, after which the furnace is cooled down to 570 °C with 10 sccm Ar carrier gas flow. The carrier gas flow is then increased to 200 sccm and the furnace opened for rapid cooling. During the growth phase, a mixture of H_2S and H_2 is introduced during the temperature ramp up phase, at different preset-temperatures in the 600 °C–700 °C range. We have tested the effect of different flow rates of the H_2S and H_2 mixture in the 5:15–15:5 sccm range. Flow rates for H_2S and H_2 are maintained constant until the start of the cool-down phase.

TEM and AFM imaging

CVD MoS_2 was transferred from sapphire using the wet transfer KOH method. Samples were first spin coated at 1500 rpm with PMMA A2, resulting in a ~100 nm thick polymer film. These were detached in a 30% KOH solution, washed several times in DI water and transferred onto TEM grids. TEM grids were annealed in a flow of Ar and H_2 for 8 h at 400 °C in order to remove the polymer film. For low-resolution imaging and diffraction studies, 10 nm thick Si_3N_4 windows were used while for high resolution TEM (HR-TEM) we used PELCO Holey Silicon Nitride Support Film with 2.5 μm holes in a 200 nm thick Si_3N_4 support. Aberration-corrected scanning transmission electron microscopy (STEM) was performed using a Titan Themis 60-300 (FEI) operated with a beam energy of 80 keV. Medium angle annular dark field imaging conditions were applied with a probe convergence semi-angle of ~28 mrad and a STEM detector collection semi-angle of ~40–200 mrad.

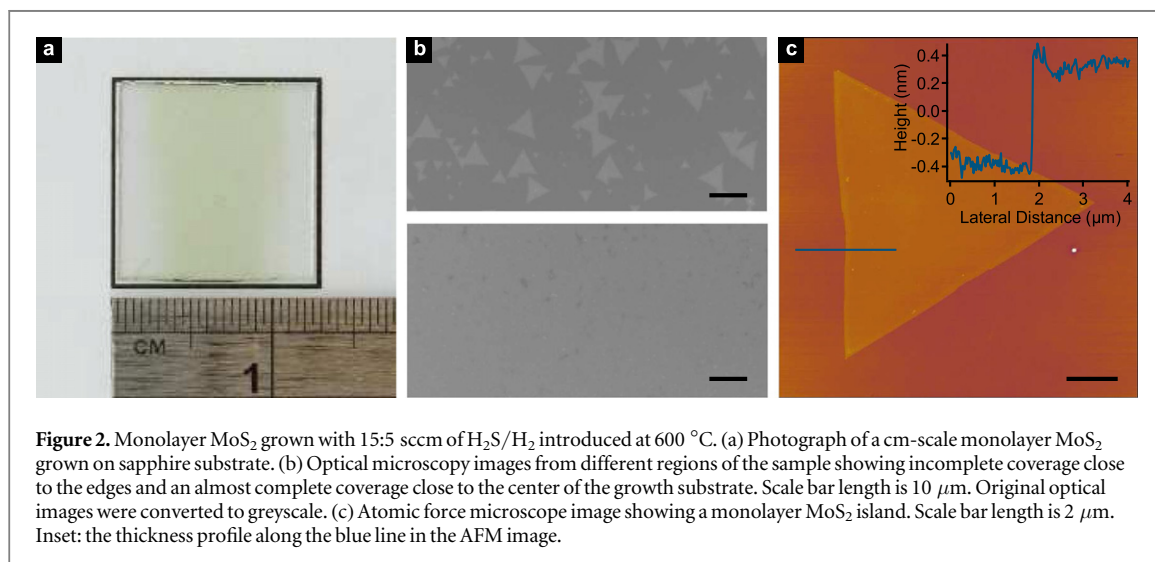
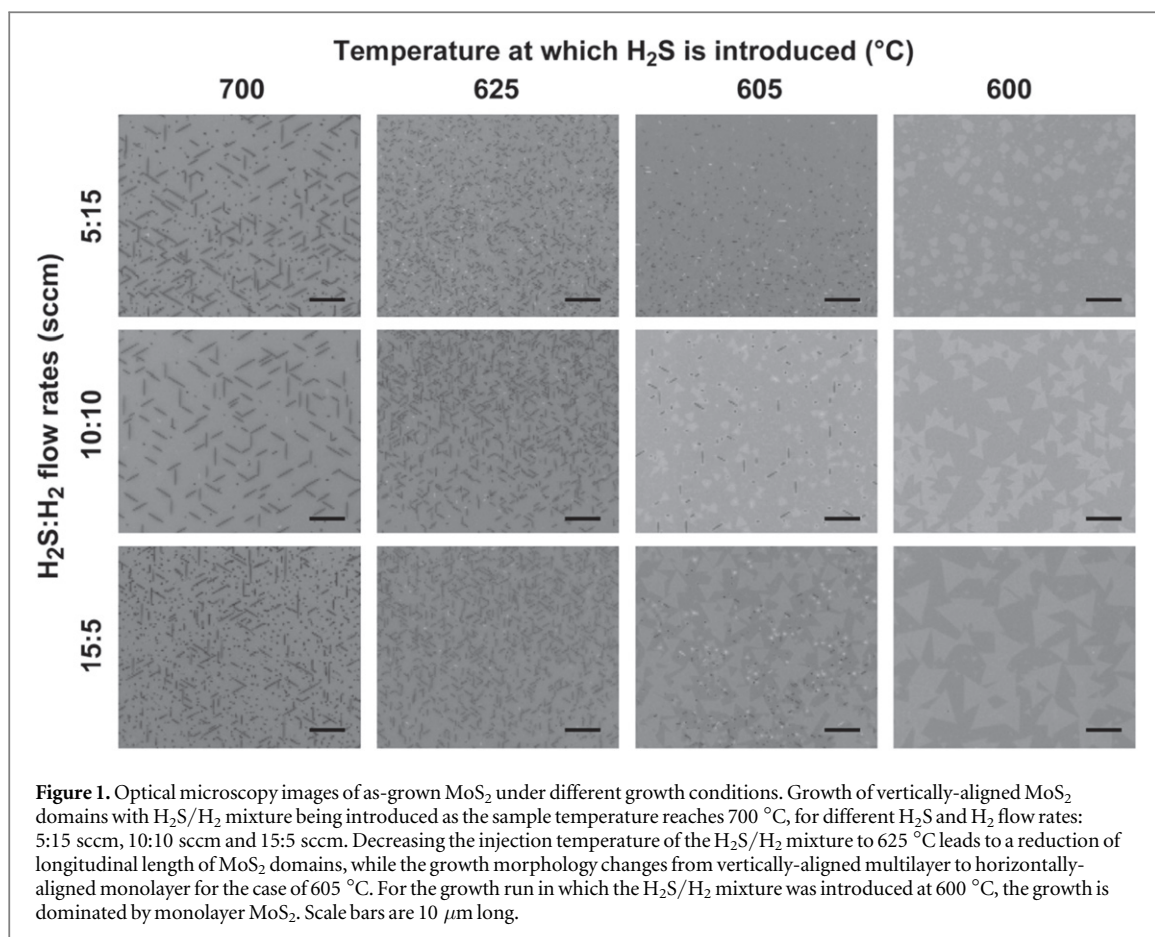
Samples were also imaged using an atomic force microscope (Asylum Research Cypher) operating in AC mode.

Electrical characterization

CVD-grown single domains of MoS_2 were transferred using PMMA A2 as a support film and etching in 30% KOH onto degenerately doped Si substrate covered with 270 nm thick SiO_2 . The PMMA film was dissolved in acetone and residues were removed by annealing in Ar atmosphere at 350 °C for 5 h. Electron beam lithography (EBL) was used to define contact regions and 90 nm Au was used as the contact metal. PMMA A4 was used as the etching mask during oxygen plasma etching for some of the devices. The devices were then annealed at 200 °C in Ar atmosphere to eliminate resist residues and reduce contact resistance. Further characterization was performed in vacuum at a base pressure of 10^{-6} mbar after annealing at 140 °C for 24 h.

Results

MoS_2 growth using H_2S as a gas-phase sulfur precursor results in highly-oriented vertically-aligned multilayers or horizontally-aligned monolayer domains in the shape of well-defined equilateral triangles (figure 1). At a high $\text{H}_2\text{S}/\text{H}_2$ injection temperature of 700 °C with different $\text{H}_2\text{S}:\text{H}_2$ ratios (5:15, 10:10 and 15:5 sccm), vertically-aligned multilayer MoS_2 domains with lengths of up to 10 μm have been observed under an optical microscope. The lateral length of MoS_2 domains decreased in size to 4 μm as the temperature at which the $\text{H}_2\text{S}/\text{H}_2$ mixture is first introduced is lowered to 625 °C. When this temperature is further reduced to 605 °C, together with vertical growth, monolayer MoS_2 domains start to grow in a lateral direction on the surface of the sapphire substrate. For an even lower injection



temperature of 600°C , the growth is dominated by monolayer MoS_2 triangles. In general, we observe that an earlier introduction of the $\text{H}_2\text{S}/\text{H}_2$ mixture favors horizontal over vertical growth. Moreover, by varying the ratio of the $\text{H}_2\text{S}/\text{H}_2$ mixture, the morphology and domain size of monolayer MoS_2 flakes is changed. Under more sulphidizing conditions ($\text{H}_2\text{S}:\text{H}_2 = 15:5$), optical images reveal that MoS_2 domains adopt a triangular shape. This is in accordance with the findings in our previous study with CVD-growth using solid sulfur as a precursor [19], and it suggests

that the triangular shape is the equilibrium form of MoS_2 domains under highly sulphurizing conditions. Under reducing ($\text{H}_2\text{S}:\text{H}_2 = 5:15$) conditions the triangular domains no longer dominate. Instead, the appearance of another type of MoS_2 structure with a hexagonal morphology has been observed. Hence, it appears that the equilibrium shape of MoS_2 domains is sensitive to the different gas compositions used in CVD growth. Such domain morphology changes have been observed before for a much higher range of $\text{H}_2\text{S}:\text{H}_2$ ratios (500 and 0.07) [37]. Furthermore, according

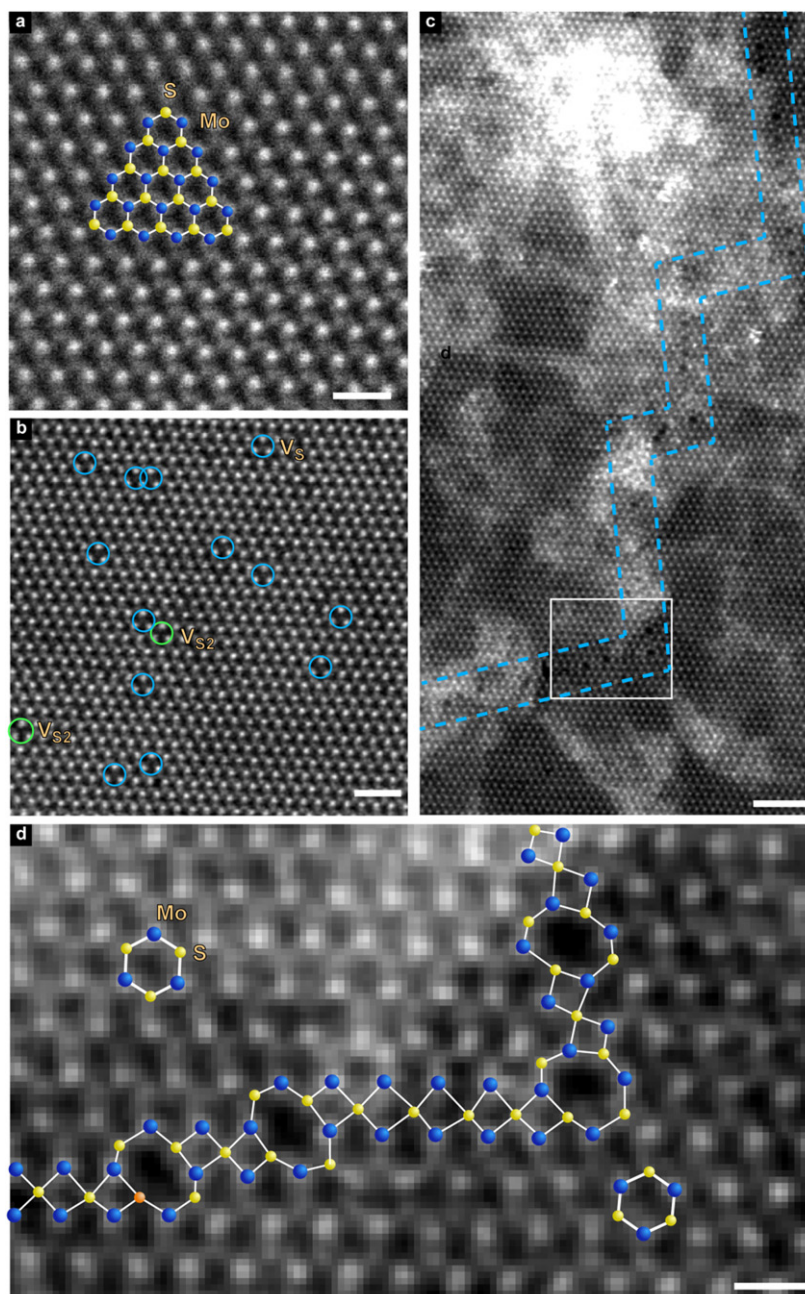
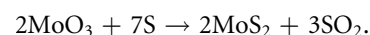


Figure 3. Monolayer MoS₂ characterization with STEM. (a) STEM image of the single-layer MoS₂ lattice with structural model overlaid. Scale bar 500 pm. (b) STEM image of a region, rich with sulfur vacancies, some of them are marked with circles. Blue: vacancies with one sulfur atom missing (V_s), green—vacancies with two sulfur atoms missing (V_{s2}). Scale bar 1 nm. (c) Overview of a grain boundary between two grains with a misorientation rotation angle close to 60°. Blue lines follow the grain boundary. Scale bar 2 nm. (d) Zoom-in view of the area marked in (c) with a white rectangle. Proposed atomic model is overlaid, blue—Mo atom, yellow—2 × S atoms, orange—1 × S atom. Scale bar 500 pm.

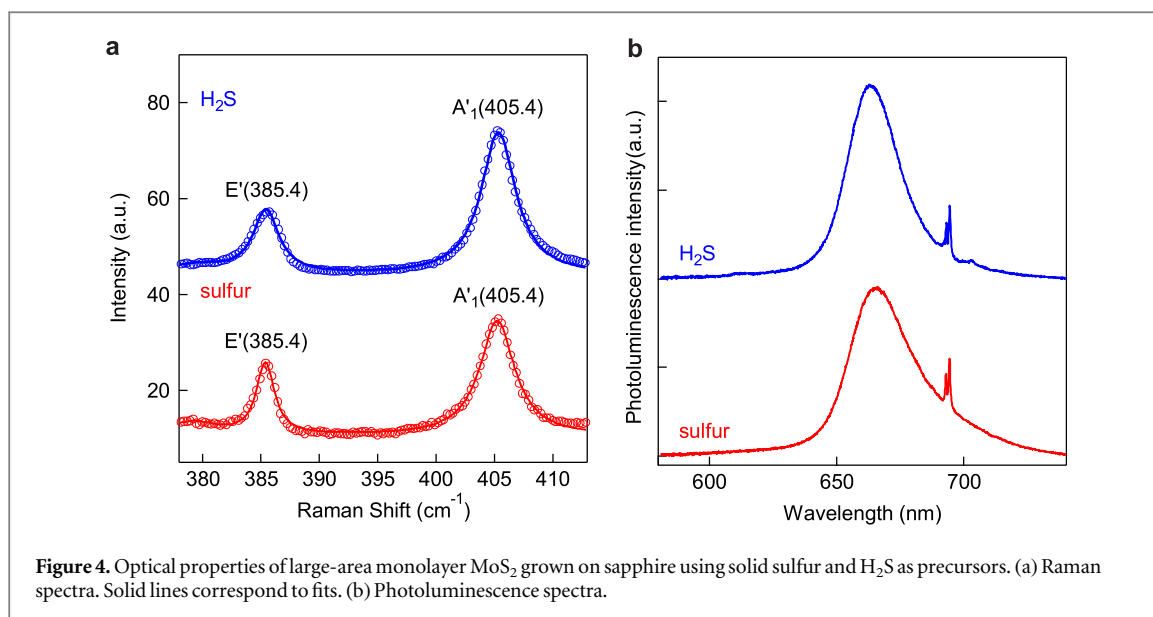
to scanning tunneling microscopy results combined with density functional theory calculations, MoS₂ triangles are terminated by dimer-saturated Mo edges, while the hexagonal MoS₂ ones exhibit completely different edge structures with lower sulfur coverage on Mo edges and S edges with adsorbed S–H groups. The presence of H atoms absorbed by S edges means that hexagonal domains could be useful in the hydrodesulfurization process [37].

We have identified optimal conditions for the growth of single-crystal domains in the shape of well-

defined equilateral triangles that merge into a continuous monolayer film, with a 15:5 sccm ratio of H₂S: H₂ gas mixture introduced at 600 °C. It is important to point out that for the syntheses of MoS₂ with solid sulfur as a precursor [19], the sulfurization of MoO₃ is typically performed at 700 °C in an Ar environment, where the chemical reaction is:



The introduction of the H₂S/H₂ gas mixture as a sulfur precursor in the reaction atmosphere leads to a



new, more favorable chemical reaction:



In this reaction, the H₂S/H₂ ratio in terms of molar volume should be equal to 2.3, which is quite close to the 15:5 ratio of H₂S:H₂ gas mixture obtained for the best growth of monolayer MoS₂. It covers a typical area of 6 mm × 1 cm in the middle portion of the growth substrate using 15:5 ratio of H₂S:H₂ gas mixture injected at 600 °C and followed by a 10 min long growth phase at 700 °C (figure 2(a)). Figures 2(b) and (c) present optical images of different sample regions, showing well separated triangles close to the substrate edge and almost full coverage of MoS₂ film in the substrate center. Figure 2(d) shows an AFM image of an as-grown triangular MoS₂ domain. The AFM line-scan demonstrates the monolayer thickness of single-crystal MoS₂ domains.

TEM characterization

To reveal the atomic structure of the material, we have performed a study of continuous film MoS₂ with aberration-corrected STEM. In order to prevent damage during initial imaging of suspended single layers, in figure 3(a) fast acquisition of several (10–15) frames with 0.5 s per frame has been used. The size of the recorded images was 512 × 512 pixels, which resulted in approximately 2–4 μs per pixel dwell time with a 0.134 Å pixel size. We used beam currents below 50 pA, which resulted in relatively low doses of $3.5 \times 10^4 \text{ e}^-/\text{Å}^2/\text{frame}$ for sub-sequential recording of each frame [38]. Finally, drift-correction and averaging of intensities with Digital Micrograph were used.

Figure 3(a) presents the lattice of MoS₂, imaged in this way and with a negligible beam-induced damage. Here, brighter spots correspond to Mo atoms and darker spots to stacks of two sulfur atoms. In general, contrast in STEM mode under these conditions of monolayer imaging is considered to obey the $Z^{1.7}$

contrast relation [39], which in the absence of contamination allows distinguishing between the atoms with different Z number or between columns with different occupancy. Extracted from Fourier-transformed STEM images and used to draw the overlaid structural model, the lattice constant of $\sim 3.18 \text{ Å}$ is in good agreement with the literature values [40]. In the single frame lower magnification images where the dose is even lower (total dose $1.25 \times 10^5 \text{ e}^-/\text{Å}^2$) one can see regions with point defects. As illustrated on figure 3(b), we can observe sulfur vacancies with one S atom absent (V_S), while occasionally we also observe vacancies with two missing S atoms (V_{S_2}), circled in green. While we cannot completely rule out that these vacancies were introduced during imaging due to beam damage, they could well be inherent to the as-grown material. Indeed, we note that similar point defects were previously observed in CVD MoS₂ [41].

Next we studied the atomic structure of grain boundaries formed by two grains with their relative orientation close to 60°. As shown on figure 3(c), the grain boundary follows a zigzag direction, and is formed mainly by 4- and 8-fold rings.

A mirror twin boundary, which equally corresponds to a 60° relative orientation, has previously been observed to show 20° misorientation zigzag boundary faceting, with a periodic 8–4–4 ring motif along the facets [18]. In our case, although the zigzag misorientation angles are different, 8-fold rings connected by two 4-fold rings are also observed, as shown on figure 3(d). In addition, a chain of five 4-fold rings that connect two differently oriented 8–4–4–8 series of rings was observed. Such a difference is most probably related to the relative translational lattice shift and the small deviation from a perfect 60° orientation between the two grains. Moreover, it has been shown that grain boundaries can consist of distinct 4-fold ring chains forming one-dimensional metallic wires [41].

However, the proposed structure of long 4-fold ring chains would require perfect angle matching and epitaxial growth of MoS₂ on sapphire [19], which was not observed in this case.

Optical properties

The quality of our monolayer MoS₂ has been checked by performing optical characterization. Figure 4(a) shows typical Raman spectra of the CVD material grown on sapphire using H₂S and solid sulfur as precursors [19]. The spectra show two characteristic peaks related to the out-of-plane vibration of S atoms (A_1') with a Raman shift of ~ 405 cm⁻¹ and the in-plane vibration of Mo and S atoms (E') at ~ 385 cm⁻¹. The observed frequency difference confirms that the deposited material is monolayer MoS₂. The position and intensity ratio between the A_1' and E' Raman modes indicate that the quality of both CVD materials is similar.

Figure 4(b) shows typical photoluminescence spectra acquired at room temperature on CVD-grown monolayer MoS₂ grown using H₂S and solid sulfur. From a sample grown by H₂S, we can clearly resolve the intense A excitonic peak positioned at 663 nm (1.87 eV) with a width of ~ 25 nm (~ 70 meV), which is similar to that of the sample grown with sulfur. The B exciton is not expected to be visible at low excitation intensities due to state-filling effects [42]. The sharp doublet at 693 and 694 nm is associated with the presence of Cr³⁺ luminescence centers in the sapphire substrate [43].

Electrical characterization

After characterization of structural and optical properties of our material, we moved to device fabrication and electrical characterization. Material was transferred from sapphire substrates to conventional Si/SiO₂ (270 nm thickness) and contacts were defined by standard EBL. Gold with a thickness of 90 nm was used as the contact material (see methods). Some of the devices were etched in rectangular shape using the second step of EBL followed by oxygen plasma etching.

We characterized our devices in vacuum (base pressure 10^{-6} mbar) after annealing at 140 °C for 24 h. We present the characteristics of a typical device after annealing on figure 5. This device was etched in rectangular shape. I_s - V_g characteristic is shown on figure 5(a), with $V_s = 1$ V applied. The field-effect mobility, extracted from two-contact measurements is presented on figure 5(b). We used the following expression to extract two-probe mobility: $\mu_{2C} = [L/W] \times [dG_{2C}/dV_g] \times [1/C_{ox}]$, where G_{2C} stands for two-probe conductivity, $C_{ox} = 1.28 \times 10^{-8}$ F cm⁻²—geometrical capacitance of 270 nm thick SiO₂, normalized per unit area. The two-probe mobility, which does not take into account the effect of contact resistance, presents the lower estimate of

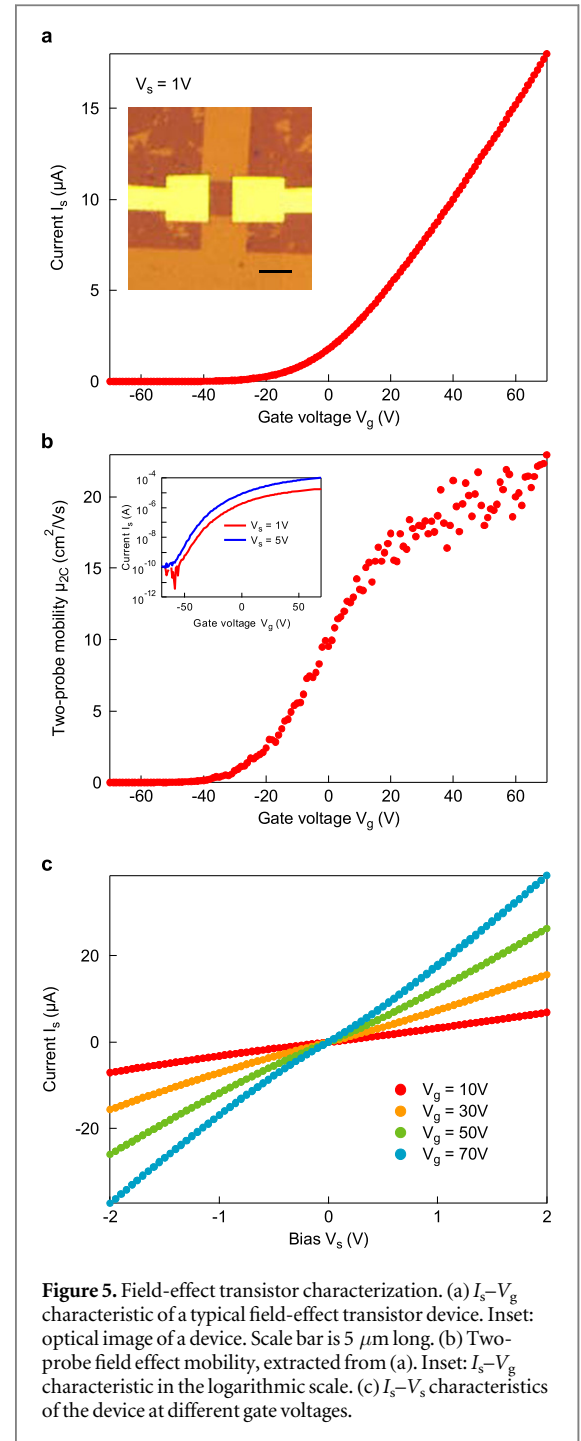


Figure 5. Field-effect transistor characterization. (a) I_s - V_g characteristic of a typical field-effect transistor device. Inset: optical image of a device. Scale bar is 5 μ m long. (b) Two-probe field effect mobility, extracted from (a). Inset: I_s - V_g characteristic in the logarithmic scale. (c) I_s - V_s characteristics of the device at different gate voltages.

intrinsic properties of the material. Values of $\mu_{2C} \sim 20$ cm² V⁻¹ s⁻¹ are in good agreement with recent reports on CVD MoS₂, grown from solid precursors on SiO₂ [44] and sapphire [19] substrates, which indicates a similar quality of material. We estimate an I_{on}/I_{off} ratio $\sim 10^6$ of this transistor at $V_s = 5$ V from the blue curve in the inset of figure 5(b). Finally, we inspect the quality of injection, performing I_s - V_s sweeps with the fixed gate voltages, shown on figure 5(c).

Overall, the device characteristics demonstrate good current injection and mobility, high I_{on}/I_{off} ratio and electron doping of the channel, typical for exfoliated or CVD material. The observed doping levels are most probably related to sulfur vacancies and other

clusters of defects, as observed in our material by STEM.

Conclusion

In conclusion, we have demonstrated controlling growth direction of MoS₂ domains by CVD using H₂S as a gas-phase precursor instead of conventional sulfur. Optical and electrical measurements on material grown in the horizontal direction, i.e. in-plane of the substrate, demonstrated the high quality of our samples, similar to previous results on exfoliated and CVD-grown samples. Aside from in-plane monolayer MoS₂, the proposed growth strategy involving the replacement of solid sulfur by H₂S gas precursor also permits one to obtain vertically grown multilayer structures that may have various electrochemical applications due to chemically reactive edge sites.

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